The World of Chemistry

Episode 13 - The Driving Forces

1. When wood burns, the system proceeds toward \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ energy and the energy of

The system \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

2. Describe the energy levels of reactants and products in an exothermic reaction.

3. In general, most chemical reactions \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ energy.

4. Why does an endothermic reaction feel cold if energy is being absorbed?

5. Describe the energy levels of reactants and products in an endothermic reaction.

6. Give two everyday examples of endothermic reactions.

7. What is meant by the term entropy?

8. In which of the following examples is entropy INCREASING?

A. Solid turns to liquid

B. Liquid turns to gas

C. Solid turns to gas

9. In general, reactions proceed toward \_\_\_\_\_\_\_\_\_\_ energy and \_\_\_\_\_\_\_\_\_\_\_\_ entropy.

10.What three factors were mentioned that determine how fast a chemical reaction takes place?